

22. Draw a sample of gas that is colder than all three of the samples in Question 21. Explain why you are sure that it is colder.

⊗ The volume ↓  
even a # of particles increase

23. Four of the relationships you investigated in Models 1 and 2 are named after scientists who discovered the relationships. Use the Internet or your textbook to match each of the scientists below with the appropriate law. Write the algebraic expression that describes the law in the box below each name.

Robert Boyle	Jacques Charles	Guillaume Amontons	Amedeo Avogadro
$P_1 V_1 = P_2 V_2$	$\frac{V_1}{T_1} = \frac{V_2}{T_2}$	$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$	$6.02 \times 10^{23}$

## Read This!

Chemists combine all of the relationships seen in Models 1 and 2 into one law—the **Ideal Gas Law**. It is one equation that describes gas behavior and the relationship among all four variables, P, V, T, and n. In the Ideal Gas Law the proportionality constant is represented by the letter R (rather than the generic k).

24. Circle the algebraic equation below that best combines all of the relationships you identified among P, V, T, and n in this activity.

$$P = RTnV$$

$$PT = RnV$$

$$PV = nRT$$

$$PTV = Rn$$

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**Nicholas R. Romano, Philip K. Cameron**

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