

Gas Laws Problem Packet

Gas Laws Packet

Name _____ Date _____ Period _____

Instructions

Standard Conditions

Unit	Standard Pressure
atm	1.013
mmHg	760
gms	1
cm	1.013

22.4 L = 1 mol at STP

STP (Standard Temperature and Pressure) = 0°C and 1 atm

The gas laws:

Boyle's Law	Charles's Law
$P_1V_1 = P_2V_2$	$\frac{V_1}{T_1} = \frac{V_2}{T_2}$
Combined Gas Law	Avogadro's Law
$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$	$\frac{V_1}{n_1} = \frac{V_2}{n_2}$
Avogadro's Law	Combined Gas Law
$\frac{V_1}{n_1} = \frac{V_2}{n_2}$	$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$
Combined Gas Law	Avogadro's Law
$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$	$\frac{V_1}{n_1} = \frac{V_2}{n_2}$
Avogadro's Law	Combined Gas Law
$\frac{V_1}{n_1} = \frac{V_2}{n_2}$	$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$

The constant R (used in the ideal gas law) depends on the unit of pressure:

- 0.0821, R = 0.0821 $\frac{\text{atm} \cdot \text{L}}{\text{mol} \cdot \text{K}}$
- 8.314, R = 8.314 $\frac{\text{J}}{\text{mol} \cdot \text{K}}$

Notes:

- Temperature must always be in Kelvin! $0^\circ\text{C} + 273 = 273\text{K}$
- If you are given two different units of pressure, convert them.
- Always use the same units for volume, and convert for the unknown. (For Ideal Gas Law, write what you are given, insert those in equation and solve for unknown)
- Avogadro's Law: DON'T FORGET TO USE MOLE! (Moles)
- Avogadro's Law: You will ALWAYS have the molecular mass for these problems. Double and triple check you have the mole of effluent and molecular mass in the correct spots before solving the equation. (This is an INVERSE relationship). You MUST calculate the square root of double both sides of the equation before moving things.

Charles' Law $\frac{V_1}{T_1} = \frac{V_2}{T_2}$

1. The temperature under my refrigerator is about 40°F. If I place a balloon in the refrigerator, will it expand or contract? Why?

2. A balloon is at 20°C and 1 atm. If I place it in a refrigerator, will it expand or contract? Why?

Boyle's Law $P_1V_1 = P_2V_2$

1. A gas is at 1 atm and 1 L. If I compress it to 0.5 L, what happens to the pressure? Why?

2. A gas is at 1 atm and 1 L. If I expand it to 2 L, what happens to the pressure? Why?

Combined Gas Law $\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$

1. A gas is at 1 atm, 1 L, and 273 K. If I change the pressure to 2 atm and the temperature to 546 K, what happens to the volume? Why?

2. A gas is at 1 atm, 1 L, and 273 K. If I change the pressure to 0.5 atm and the temperature to 546 K, what happens to the volume? Why?

Avogadro's Law $\frac{V_1}{n_1} = \frac{V_2}{n_2}$

1. A gas is at 1 atm, 1 L, and 273 K. If I add 1 mole of gas, what happens to the volume? Why?

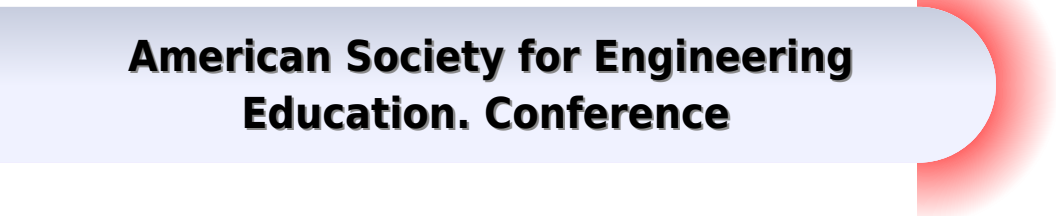
2. A gas is at 1 atm, 1 L, and 273 K. If I remove 1 mole of gas, what happens to the volume? Why?

All Gas Laws Included!



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