

## CHAPTER 21 REVIEW

***Nuclear Chemistry*****SECTION 1****SHORT ANSWER** Answer the following questions in the space provided.

1. \_\_\_\_\_ Based on the information about the three elementary particles in the text, which has the greatest mass?
  - (a) the proton
  - (b) the neutron
  - (c) the electron
  - (d) They all have the same mass.
2. \_\_\_\_\_ The force that keeps nucleons together is
  - (a) a strong nuclear force.
  - (b) a weak nuclear force.
  - (c) an electromagnetic force.
  - (d) a gravitational force.
3. \_\_\_\_\_ The stability of a nucleus is most affected by the
  - (a) number of neutrons.
  - (b) number of protons.
  - (c) number of electrons.
  - (d) ratio of neutrons to protons.
4. \_\_\_\_\_ If an atom should form from its constituent particles,
  - (a) matter is lost and energy is taken in.
  - (b) matter is lost and energy is released.
  - (c) matter is gained and energy is taken in.
  - (d) matter is gained and energy is released.
5. \_\_\_\_\_ For atoms of a given mass number, those with greater mass defects, have
  - (a) smaller binding energies per nucleon.
  - (b) greater binding energies per nucleon.
  - (c) the same binding energies per nucleon as those with smaller mass defects.
  - (d) variable binding energies per nucleon.
6. Based on **Figure 1.1** of the text, which isotope of He, helium-3 or helium-4,
  - \_\_\_\_\_ a. has the smaller binding energy per nucleon?
  - \_\_\_\_\_ b. is more stable to nuclear changes?
7. The number of neutrons in an atom of magnesium-25 is \_\_\_\_\_.
8. Nuclides of the same element have the same \_\_\_\_\_.

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## Answer Key

**Martin Silberberg**



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## **Table of Contents Glencoe Science Chemistry Chapter 23 Study Guide Answer Key**

1. Understanding the eBook Glencoe Science Chemistry Chapter 23 Study Guide Answer Key
  - The Rise of Digital Reading Glencoe Science Chemistry Chapter 23 Study Guide Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Glencoe Science Chemistry Chapter 23 Study Guide Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Glencoe Science Chemistry Chapter 23 Study Guide Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Glencoe Science Chemistry Chapter 23 Study Guide Answer Key
  - Personalized Recommendations

- Glencoe Science Chemistry Chapter 23 Study Guide Answer Key User Reviews and Ratings
- Glencoe Science Chemistry Chapter 23 Study Guide Answer Key and Bestseller Lists
- 5. Accessing Glencoe Science Chemistry Chapter 23 Study Guide Answer Key Free and Paid eBooks
  - Glencoe Science Chemistry Chapter 23 Study Guide Answer Key Public Domain eBooks
  - Glencoe Science Chemistry Chapter 23 Study Guide Answer Key eBook Subscription Services
  - Glencoe Science Chemistry Chapter 23 Study Guide Answer Key Budget-Friendly Options
- 6. Navigating Glencoe Science Chemistry Chapter 23 Study Guide Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Glencoe Science Chemistry Chapter 23 Study Guide Answer Key Compatibility with Devices
  - Glencoe Science Chemistry Chapter 23 Study Guide Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Glencoe Science Chemistry Chapter 23 Study Guide Answer Key
  - Highlighting and Note-Taking Glencoe Science Chemistry Chapter 23 Study Guide Answer Key
  - Interactive Elements Glencoe Science Chemistry Chapter 23 Study Guide Answer Key
- 8. Staying Engaged with Glencoe Science Chemistry Chapter 23 Study Guide Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Glencoe Science Chemistry Chapter 23 Study Guide Answer Key
- 9. Balancing eBooks and Physical Books Glencoe Science Chemistry Chapter 23 Study Guide Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Glencoe Science Chemistry Chapter 23 Study Guide Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Glencoe Science Chemistry Chapter 23 Study Guide Answer Key
  - Setting Reading Goals Glencoe Science Chemistry Chapter 23 Study Guide Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Glencoe Science Chemistry Chapter 23 Study Guide Answer Key
  - Fact-Checking eBook Content of Glencoe Science Chemistry Chapter 23 Study Guide Answer Key

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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