

Chemistry Study Guide

Periodic Table of the Elements

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| <small>3</small> Li <small>6.941</small> | <small>4</small> Be <small>9.012</small> | | | | | | | | | | | | | | | | | <small>5</small> B <small>10.81</small> | <small>6</small> C <small>12.01</small> | <small>7</small> N <small>14.01</small> | <small>8</small> O <small>16.00</small> | <small>9</small> F <small>18.998</small> | <small>10</small> Ne <small>20.18</small> | | | | | | | | | | | | | | | | | | | | | | | | | |
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Objective:

To understand the *structure of an atom*, and the *nature + properties* of subatomic particles ✎

Matter:

Anything that has **mass** and **occupies space** ✎

Examples: air, gases, liquids, solids!

Atoms:

Atoms are the **smallest, indivisible particle** ✎

- **Outer part** → *shell/orbit/energy level* → carries the electrons (e^-)
- **Inner part** → *nucleus* → houses the protons (p^+) and neutrons (n)
- Outer & Inner part → together these are the **subatomic particles**

Subatomic Particle Table:

| Particle | Symbol | Mass (amu) | Charge | Location |
|----------|--------|------------|--------|-------------|
| Proton | p^+ | 1 | +1 | Nucleus |
| Neutron | n | 1 | 0 | Nucleus |
| Electron | e^- | ~ 0 | -1 | Orbit/Shell |

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